

Laily B. Din, Intan S. C. Sulaiman  
and Bohari M. Yamin\*School of Chemical Sciences and Food  
Technology, Universiti Kebangsaan Malaysia,  
43600 Bangi, Selangor, MalaysiaCorrespondence e-mail:  
bohari@pkrisc.cc.ukm.my

## Key indicators

Single-crystal X-ray study  
 $T = 298\text{ K}$   
Mean  $\sigma(\text{C}-\text{C}) = 0.005\text{ \AA}$   
Disorder in main residue  
 $R$  factor = 0.048  
 $wR$  factor = 0.106  
Data-to-parameter ratio = 8.5For details of how these key indicators were  
automatically derived from the article, see  
<http://journals.iucr.org/e>.

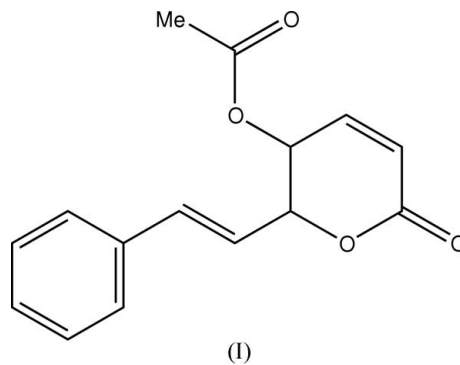
## 5-Acetylgoniothalamine

In the title compound,  $\text{C}_{13}\text{H}_{10}\text{O}_2$ , the pyrone ring adopts a modified screw-boat conformation and a *trans* configuration about the  $\text{C}=\text{C}$  double bond. The styryl fragment is pseudo-equatorial on the pyrone ring. The molecules are stabilized by intermolecular  $\text{C}-\text{H}\cdots\text{O}$  interactions and form a one-dimensional zigzag chain parallel to  $[100]$ .

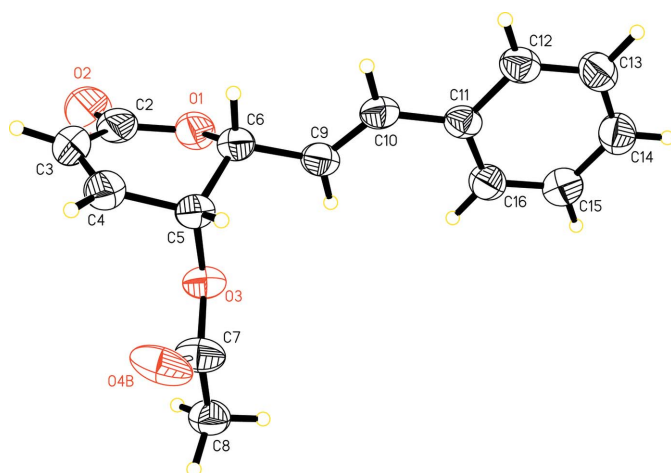
Received 30 June 2006  
Accepted 10 July 2006

## Comment

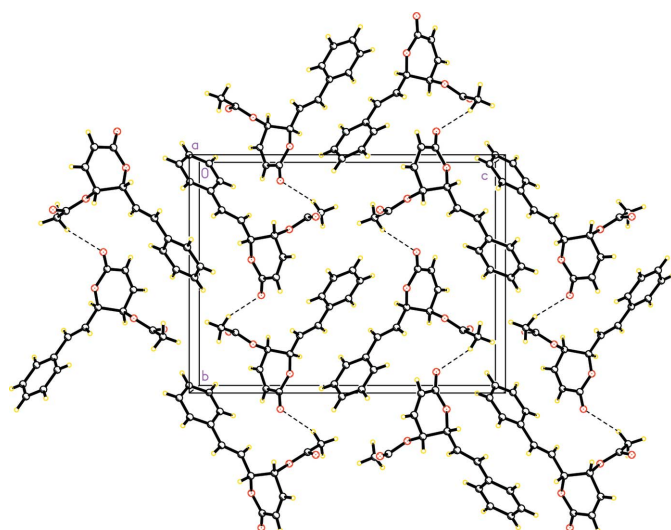
Biologically active styrylpyrone derivatives such as (+)-goniothalamine (Jewers *et al.*, 1972) and (+)-goniothalamine epoxide (Sam *et al.*, 1987) are among the dominant compounds originally isolated from the genus *Cryptocarya* (Hlubucek & Robertson, 1967) and '*Goniothalamus*' (Airy-Shaw, 1966). 5-Acetyl goniothalamine (Fig. 1), a styryl-dihydropyrone, was first reported as the second major component after goniothalamine, isolated from the species *G. uvaroides* King, collected from Sabah, Malaysia (Ahmad *et al.*, 1991).



In the present structural investigation, the title compound, (I), was isolated from the same genus but of species *G. ridleyi* King. The molecular structure of (I) is typical of  $\alpha\beta$ -unsaturated  $\delta$ -lactones and similar to 5,6-dihydro-6-styryl-2-pyrone, (II) (Clarke & Pauling, 1975). The replacement of an H atom by the acetyl group at atom C5 of the pyrone ring showed no significant effect on the bond lengths of the ring except that the  $\text{C4}-\text{C5}-\text{C6}$ ,  $\text{C3}-\text{C4}-\text{C5}$  and  $\text{C4}-\text{C3}-\text{C2}$  bond angles (Table 1) are slightly smaller than those in (II) [ $110.7(4)$ ,  $120.6(5)$  and  $122.7(4)^\circ$ , respectively]. The conformation of the ring is a modified screw-boat as indicated by the non-planarity of the ring, with a maximum deviation of  $0.328(3)\text{ \AA}$  for atom C6 from the least-squares plane of the  $\text{O1}/\text{C2}-\text{C6}$  fragment. Atom C3 deviates by  $0.161(3)\text{ \AA}$  on the same side as atom C6. Atoms C5 [deviation  $0.275(3)\text{ \AA}$ ], O1 [deviation  $0.127(2)\text{ \AA}$ ] and C2 [deviation  $0.115(2)\text{ \AA}$ ] are on the oppo-



**Figure 1**  
Molecular structure of (I), with 50% probability displacement ellipsoids. The minor component of the disorder has been omitted.



**Figure 2**  
Packing diagram of the title complex, viewed down the *a* axis. The dashed lines denote C—H...O hydrogen bonds.

site side of the least-squares plane. The strain-relieving factor of the  $\alpha\beta$ -unsaturated  $\delta$ -lactone indicated by the torsion angle O2—C2—C3—C4 of 157.3 (4) $^\circ$  is less than that in (II) [169.3 (6) $^\circ$ ]. The molecule maintains its *trans* configuration about the C9=C10 bond and the pseudo-equatorial position of the planar styryl ring C9—C16 [maximum deviation 0.022 (2) Å for C10 atom] on the pyrone ring. The two rings are approximately perpendicular. This is in contrast to dehydrogoniothalamin (Din *et al.*, 2004), where the whole molecule is essentially planar. The crystal structure is stabilized by weak C—H...O molecular interactions (Table 2), forming zigzag polymeric chains parallel to the *b* axis (Fig. 2).

## Experimental

*Goniothalamus ridleyi* King was collected from Gua Musang, Kelantan, East of Peninsular Malaysia. Stem bark powder (420 g) was extracted in petroleum ether (2 l) in a Soxhlet apparatus for 36 h. Yellow crystals were obtained at the bottom of the Soxhlet flask. The

crystals were washed with petroleum ether and recrystallized from ethyl acetate to afford colourless crystals (1.2 g, 0.3%), with melting point 391–392 K.  $^1\text{H}$  NMR ( $\text{CDCl}_3$ , p.p.m.):  $\delta$  7.26 (*m*, 5 H, aromatic), 6.90 (*dd*,  $J = 10$  and 5 Hz, H-4), 6.80 (*dd*,  $J = 15$  and 1 Hz, H-8), 6.20 (*dd*,  $J = 10$  and 2 Hz, H-3), 6.16 (*dd*,  $J = 15$  and 6 Hz, H-7), 5.26 (*dd*,  $J = 6$  and 3 Hz, H-6), 5.15 (*dd*,  $J = 5$  and 3 Hz, H-5) and 2.02 (*s*, 3 H, —OMe).  $^{13}\text{C}$  NMR ( $\text{CDCl}_3$ , p.p.m.):  $\delta$  170.2 (*s*, ester carbonyl), 162.6 (*s*, C-2), 140.8 (*d*, C-4), 135.7 (*d*, C-9), 135.0 (*d*, C-7), 128.8 (*d*, C-11 dan C-13), 128.7 (*d*, C-12), 126.9 (*d*, C-10 and C-14), 125.0 (*d*, C-8), 121.2 (*d*, C-3), 79.2 (*d*, C-6), 64.0 (*d*, C-5) and 20.7 (*q*, Ac).

## Crystal data

$\text{C}_{15}\text{H}_{14}\text{O}_4$	$Z = 4$
$M_r = 258.26$	$D_x = 1.286 \text{ Mg m}^{-3}$
Orthorhombic, $P2_12_12_1$	Mo $K\alpha$ radiation
$a = 5.3798$ (12) Å	$\mu = 0.09 \text{ mm}^{-1}$
$b = 13.659$ (3) Å	$T = 298$ (2) K
$c = 18.153$ (4) Å	Block, colourless
$V = 1333.9$ (5) Å $^3$	$0.49 \times 0.33 \times 0.18 \text{ mm}$

## Data collection

Bruker SMART APEX CCD area-detector diffractometer	7462 measured reflections
$\omega$ scans	1551 independent reflections
Absorption correction: multi-scan (SADABS; Bruker, 2000)	1234 reflections with $I > 2\sigma(I)$
$T_{\min} = 0.955$ , $T_{\max} = 0.983$	$R_{\text{int}} = 0.025$
	$\theta_{\max} = 26.0^\circ$

## Refinement

Refinement on $F^2$	$w = 1/[\sigma^2(F_o^2) + (0.0402P)^2 + 0.207P]$
$R[F^2 > 2\sigma(F^2)] = 0.048$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.106$	$(\Delta\sigma)_{\max} < 0.001$
$S = 1.12$	$\Delta\rho_{\max} = 0.12 \text{ e } \text{Å}^{-3}$
1551 reflections	$\Delta\rho_{\min} = -0.12 \text{ e } \text{Å}^{-3}$
182 parameters	
H-atom parameters constrained	

**Table 1**

Selected geometric parameters (Å,  $^\circ$ ).

O1—C2	1.347 (4)	C4—C5	1.487 (4)
O1—C6	1.459 (4)	C5—C6	1.505 (4)
O2—C2	1.201 (5)	C6—C9	1.484 (4)
C2—C3	1.469 (6)	C7—C8	1.457 (5)
C3—C4	1.308 (5)	C9—C10	1.312 (4)
C4—C3—C2	121.3 (3)	C4—C5—C6	108.6 (3)
C3—C4—C5	119.8 (4)	C2—C3—C4—C5	6.3 (6)
C6—O1—C2—O2	174.5 (3)	C3—C4—C5—C6	30.1 (5)
C6—O1—C2—C3	−9.1 (4)	O1—C2—C3—C4	−54.3 (4)
O1—C2—C3—C4	−18.9 (6)		

**Table 2**

Hydrogen-bond geometry (Å,  $^\circ$ ).

$D\text{—H}\cdots A$	$D\text{—H}$	$\text{H}\cdots A$	$D\cdots A$	$D\text{—H}\cdots A$
C8—H8A...O4B <sup>i</sup>	0.96	2.59	3.120 (8)	115
C8—H8B...O4B <sup>ii</sup>	0.96	2.59	3.382 (6)	140
C8—H8C...O2 <sup>iii</sup>	0.96	2.43	3.315 (4)	153

Symmetry codes: (i)  $x - 1, y, z$ ; (ii)  $x - \frac{1}{2}, -y + \frac{1}{2}, -z + 1$ ; (iii)  $-x, y + \frac{1}{2}, -z + \frac{3}{2}$ .

Atom O4 is disordered and the site-occupation factors of the disorder components O4A and O4B refined to 0.49 (2) and 0.51 (2), respectively. In the absence of significant anomalous scattering, Friedel pairs were averaged. H atoms were positioned geometrically with C–H = 0.93–0.98 Å, and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C})$ , where  $x = 1.5$  for methyl H and  $x = 1.2$  for all other H.

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINTE* (Bruker, 2000); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997a); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997a); molecular graphics: *SHELXTL* (Sheldrick, 1997b); software used to prepare material for publication: *SHELXTL*, *PARST* (Nardelli, 1995) and *PLATON* (Spek, 2003).

The authors thank the Malaysian Government and Kebangsaan Malaysia for the research grants IRPA Nos. 09-02-03-0145 and 09-02-02-0163, respectively.

## References

- Ahmad, F. B., Tukoi, W. A., Omar, S. & Sharif, A. M. (1991). *Phytochemistry*, **30**, 2430–2431.
- Airy-Shaw, H. K. (1966). *A Dictionary of Flowering Plants and Ferns*, p. 489. Cambridge University Press.
- Bruker (2000). *SADABS* (Version 2.01), *SMART* (Version 5.630a) and *SAINTE* (6.63a). Bruker AXS Inc., Madison, Wisconsin, USA.
- Clarke, P. J. & Pauling, P. J. (1975). *J. Chem. Soc. Perkin Trans. 2*, pp. 368–370.
- Din, L. B., Abdullah, A., Tan, S. L., Said, I. M. & Yamin, B. M. (2004). *Acta Cryst. E* **60**, o1515–o1517.
- Hlubucek, J. R. & Robertson, A. V. (1967). *Aust. J. Chem.* **20**, 2199–2206.
- Jewers, K., Davis, J. B., Dougan, J., Manchanda, A. H., Blunden, G., Kyi, A. & Wetchapinan, S. (1972). *Phytochemistry*, **11**, 2025–2030.
- Nardelli, M. (1995). *J. Appl. Cryst.* **28**, 659.
- Sam, T. W., Yeu, C. S., Matsjeh, S., Gan, E. K., Razak, D. & Mohamed, A. L. (1987). *Tetrahedron Lett.* **28**, 2541–2544.
- Sheldrick, G. M. (1997a). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Sheldrick, G. M. (1997b). *SHELXTL*. Version 5.1. Bruker AXS Inc., Madison, Wisconsin, USA.
- Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.